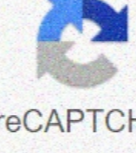


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- 4. Covalent or molecular substances have lower melting and boiling points, do not conduct electricity.
- 4. Polar substances are dissolved only by another polar substance. Non-polar substances are dissolved only by other non-polar substances.
- 4. Chemical bonds are formed when valence electrons are:
 - 4. Transferred from one atom to another - ionic.
 - 4. Shared between atoms - covalent.
 - 4. Made in a free sharing "sea" of electrons - metallic.
- 5. In multiple (double or triple) covalent bonds more than 1 pair of electrons are shared between two atoms.
- 6. Polarity of a molecule can be determined by its shape and the distribution of the charge.
 - 4. Polar molecules must have polar bonds.
 - 4. Polar molecules are asymmetrical.
 - 4. Nonpolar molecules are symmetrical and/or have no polar bonds.
- 7. When an atom gains an electron, it becomes a negative ion and its radius increases.
- 8. When an atom loses an electron, it becomes a positive ion and its radius decreases.
- 9. Atoms gain a stable electron configuration by bonding with other atoms.
 - 4. Atoms are stable when they have a full valence level.
 - 4. Most atoms need 8 electrons to fill their valence level.
 - 4. H and He only need 2 electrons to fill their valence level.
 - 4. The noble gases (group 18) have filled valence levels. They do not normally bond with other atoms.

- Chapter 2: Structure of Matter (11.1-11.3, 11.5-11.6)
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